

New directions for carbon-based detectors: exploiting the versatility of carbon substrates in electroanalysis

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Received: 17 February 2008 / Accepted: 19 February 2008 / Published online: 28 March 2008
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Abstract The different approaches that have been taken in the development of electrochemical detectors incorporating carbon as the principal electrode substrate are reviewed. The multitude of forms that carbon can take provides an extremely versatile platform upon which to enhance the detector performance. The different approaches that have been taken in recent years have been critically appraised and the emergence of new composite technologies is highlighted. The role of carbon and the applications to which it has been employed are presented and compared and contrasted within the remit of electroanalytical investigations.

Keywords Chromatography · LC-ED · FIA · Detectors · Electroanalysis

Abbreviations

GC	glassy carbon
CF	carbon fibre
RVC	reticulated vitreous carbon
HOPG	highly orientated pyrolytic graphite
eppg	edge plane pyrolytic graphite
bppg	basal plane pyrolytic graphite
CNTs	carbon nanotubes
CMEs	chemically modified electrodes
SWCNT	single-wall carbon nanotube
MWCNT	multi-walled carbon nanotube
SPE	screen-printed electrode
LC	liquid chromatography
Cyt C	cytochrome C

OPP	over-oxidised polypyrrole
GOD	glucose oxidase
HAO	histamine oxidase
ChOD	choline oxidase
GlutaOD	glutamate oxidase
NOX	NADH oxidase
GluDH	glutamate dehydrogenase
XOD	xanthine oxidase
PB	Prussian blue
MPT	methyl-patanthion
AA	ascorbic acid
HRP	horseradish peroxidase
PQQ	pyrrolquinoline quinone
IgG	immunoglobulin G
N/A	not applicable

Introduction

The combination of electrochemical detection with liquid chromatographic techniques (LC-ED) has proven to be one of the more versatile analytical options when attempting to characterise complex matrices. The detectors have taken many forms and have exhibited sensitivities that can compliment or in many instances outperform conventional ultraviolet (uv) spectroscopic techniques [1–14]. They can offer greater procedural simplicity than fluorimetric systems [15–19] and are significantly less expensive, both to procure and operate, than mass spectroscopic detection (LC-MS) [20–24]. While the sensitivities, resolute capacity and structure determination capabilities of the latter procedure are undeniably impressive, LC-ED techniques have sought to fulfil more the demands required by routine determinations without compromising selectivity. Signifi-

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cant improvements in performance have been realised through the modification of the detector surfaces to confer greater sensitivity in particular applications [1–3, 25]. The aim of the present review is to provide a critical evaluation of the roles that carbon can play in the construction of LC detectors and to give some insights into present and future applications.

While carbon has long been recognised as a versatile substrate in electroanalysis—its use in liquid chromatographic systems has been relatively understated. The emphasis of the present investigation is solely on the post-column analysis and detector design—construction highlighting the potential opportunities that carbon—in its many varied forms—can provide for LC-ED systems. The various modifications that can be undertaken to enhance detector performance and the emergence and potential impact of new composite materials are considered in the following sections.

Core detector design

Electrochemical detectors generally take one of three core design formats and relate principally to how the working electrode is arranged with respect to the direction of the mobile phase flow. Schematics outlining these are detailed in Fig. 1. The simplest relates to the insertion of a planar electrode within the flow system operation. This can be extended to porous carbon systems (typically pyrolytic graphite) that encompass the entire diameter of the flow channel. The latter systems are designed more for coulometric analysis in which the complete conversion of the analyte can be achieved [1, 2]. The third design inserts the electrode perpendicular to the flow (wall jet). Each methodology has its merits depending on the nature of the electrode being used and there will always be variations in the design of the cell used within research environments. The construction of arrays introduces an extra dimension to those already outlined in Fig. 1 and can greatly enhance the

selectivity allowing post-column resolution at the detector. Multiple detectors placed in series, each set at a different operating potential, can offer multi-dimensional analysis and is analogous to the diode array detectors though at considerably less expense. Hence, two or more analytes eluting at same time can be resolved at the detection array as a consequence of their potential dependent responses at each electrode. Thus, in a recent example, a single chromatographic signal arising from a conventional uv-visible detector was subsequently shown to possess more than 300 discrete components when resolved using a four-channel coulometric detector [2].

It has long been recognised that different electrode substrates (even when poised at identical operating potentials) can elicit significantly different responses from a given analyte. As a result, post-column electrochemical detectors have exploited a variety of electrode substrates (Pt, Au, Ag, C, Cu, Ni) and modified systems (alloys, organometallic catalysts, enzymes) to improve analytical performance [1, 3]. Carbon can be viewed as the most flexible of these substrates as a consequence of the variety of physical forms it can take but each of these can come with a diverse range of chemical properties (or functionalities) that can allow detector enhancements through a surface modification. In this respect, carbon is a significantly more flexible diagnostic tool.

Carbon: a versatile substrate

The multitude of physical forms that carbon can take allows for considerable flexibility in the design of the detector. Glassy carbon (GC), carbon aerogels, carbon fibre (CF), carbon felts, reticulated vitreous carbon (RVC) or graphite have all been used in a great number of analytical techniques [25]. Highly orientated pyrolytic graphite (HOPG) offering edge plane (eppg) and basal plane (bppg) morphologies are frequently used to explore the more fundamental side of electron transfer processes but can be of considerable value in LC-ED systems. The basal and edge plane possess different properties with the latter morphology tending to exhibit considerably faster electrode kinetics and, as a consequence, possesses the potential for greater detection sensitivities [26–29]. The edge plane sites on carbon nanotubes (CNTs) are also thought to be a major contributor to the high electron transfer rates and electrocatalytic effects observed in such systems [26, 29]. Glassy carbon is however the most commonly employed electrode material in electroanalysis and is easily adapted for use in the detector designs outlined in Fig. 1. Different studies reporting the use of carbon electrodes coupled with flow systems are presented in Table 1 and highlight the diversity of application to which carbon could be utilised and

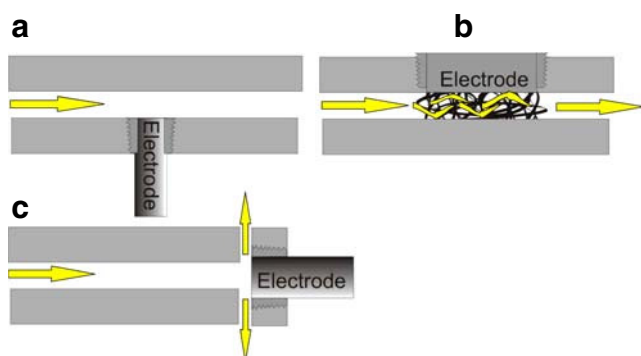


Fig. 1 Common detector designs

Table 1 Different applications of carbon electrodes

Form	Analyte	Application	Detection limits (μM)	Ref
GC	Rocuronium impurities	Pharmaceutical	0.094–235	[9]
Carbon fibre micro-disk (CFE)	Herb <i>Acanthopanax senticosus</i>	Environmental	10–1.5	[10]
GC	Ag-ions	Biochemistry	20–67	[11]
eppg	Thiols	Bioanalytical	1.7–2.7	[27]
eppg	Sulphide	Analytical	4.9–10	[28]

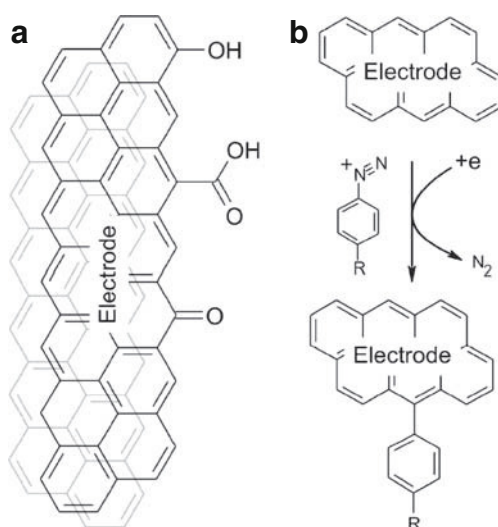
provides an indication of the detection limits than can be achieved at the unmodified electrode surfaces.

Surface modification

The surface of carbon can be modified by mechanical [28–30], electrochemical [31–38], chemical [38–45] or thermal [46] means—the aim being to alter the interfacial reactivity of the base substrate and thereby improve the detector performance. In this review, surface modification is categorised in three levels. The first level is simply a modification of the species already present on the surface (Fig. 2a); on the second level—the modification implies the attachment of exogenous species onto the electrode surface (Fig. 2b); and in the last level, the species are intercalated between the graphene layers of the base substrate (not shown). The latter level can be induced through the electrochemical pre-treatment processes as a result of carbon fracturing [31, 36, 41]. The activation of the electrode is a term that is frequently used to describe a variety of superficial processes: the elimination of electrode's impurities, the creation of a fresh electrode surface, the increase of surface functional groups or an increase in

electrode area. Mechanical activation such as polishing or sonication is widely used to clean the electrode and to expose fresh electrode area but can have the beneficial effect of increasing the availability of edge sites [27, 29] that can, as with eppg electrodes, serve to increase sensitivity. The most usual activation technique involves the electro-oxidation of the carbon surface by applying a fixed anodic or cyclic potential. The oxidation increases the proportion of oxygen species (Fig. 2a) such as hydroxyl, carbonyl, carboxyl and quinones at the electrode surface [38]. These are well known to influence the wettability, surface reactivity, porosity and conductivity of the electrode [25, 31, 32, 34].

Chemically modified electrodes signify a more elaborate approach to surface modification and generally involve the deliberate attachment of new chemical components. This offers greater control over the electrochemical behaviour given the greater variety of functionality that can be introduced onto the surface—more so than that obtained by the activation of the intrinsic functionalities described previously. Three general approaches can be taken—formation of covalent bonds [42–45] and irreversible adsorption [41]. Electrochemically induced functionalisation of surfaces by exogenous species can also be utilised and provides an effective alternative to purely chemical methods. A more recent strategy involves the reduction of diazo species (shown in Fig. 2b) whereby there is the direct covalent attachment of the target species [36, 37, 41, 43]. The most commonly adopted procedure is however immobilisation of chemical species in a polymer film onto the carbon surface [46–91] and this will be considered in more detail in the following section. Irrespective of the procedure, the modifications can be used to good effect to influence hydrophilicity–hydrophobicity of the interface and thereby alter the adsorption characteristics of the matrix components—facilitating the accumulation of one species over another and thereby improving electroanalytical discrimination [25, 30, 31, 32, 34]. Alternatively, electrocatalytically active molecules or mediators bonded to the electrode surface can act as electron transfer bridges between the electrode and electroactive species in solution [55, 60, 70] and thereby modify the electrode properties of

**Fig. 2** a Carbon functionalities and b surface activation

the base electrode. Some of the possible surface modifications and their application are detailed in Table 2.

Polymer coatings

The use of polymers as a coating material has been used extensively to enhance the electrode performance and represents the most common approach to electrode modification [73]. The chemical and physical properties of the polymer will substantially influence the electrode response characteristics with the underlying carbon normally acting as a support material and the electrical conduit. In many cases, the form of carbon utilised will be determined more by availability and by detector design considerations. This is not, however, to suggest that the inherent electrochemical properties of the base material can simply be ignored as it is still important to establish fast electrode kinetics [28, 30] between the substrate and the electroactive material within the polymer. The dimensions (micro, macro, fibre), porosity, mechanical integrity and processing capability of the base substrate and the ease with which it can be incorporated within a flow cell will often be considerable factors in selecting the base substrate. The versatility of carbon in this respect contrasts the use of metal electrodes as the detecting substrate where the physical forms available are more restrictive. Glassy carbon, reticulated vitreous carbon [25] and pyrolytic graphite [30] are the most common electrodes in comparison to other carbon materials though polymer encapsulated carbon fibre systems have grown in popularity in recent years [50–53]. The latter carbon material can be rationalised in terms of the

greater dimensional compatibility of the fibre systems within conventional LC architectures.

Numerous polymeric systems have been investigated and can be placed within several broad categories: non-conducting [54–63], conducting [64–83] and redox gels [84–88]. While a large variety of systems have been examined—the most common and arguably the most successful tend to be those based on Nafion® [59, 60], polyaniline derivatives [64–74] and polypyrrole [51, 75–83]. The first falls across the pre-formed and non-conducting categories of polymer modification and possess important anionic properties that can confer significant ion exchange properties to amperometric detectors. This is also true of the pyrrole and aniline systems but these possess the added advantage of being formed through electropolymerisation and hence offer greater control over polymer morphology (thickness, permeability etc.) and also the spatial localisation of the film. The latter advantage is important given the increasing interest in miniaturised systems where the conventional solvent casting techniques common to Nafion® would not be appropriate. Some of the applications in which polymer-modified carbon electrodes play a central role are detailed in Table 3.

The conducting polymers, especially those based upon pyrrole, can serve as an extremely flexible diagnostic tool offering a number of roles—electrocatalysis, molecular filter, enzyme entrapment etc. [51, 81–83]. Polypyrrole can be over-oxidised resulting in a loss of conductivity and the expulsion of the dopant ions [81–83, 69]. In such instances, it has been shown to possess cation exchange (as opposed to the anion exchange when in the conductive

Table 2 Surface-modified electrode and their application

Form	Analyte	Pre-treatment–modification	Application	Detection limit (μM)	Ref
GC	5-Hydroxyindoleacetyl 5-Hydroxytryptamine	Electrochemical oxidation H_2SO_4 Gold adsorption	Urine	0.0005 0.00025	[13]
GC	H^+	Electrochemical oxidation (NH_4)	pH sensor	N/A	[31]
GC	Vitamins B ₂ , B ₆ , C	Electrochemical oxidation phosphate-buffered salt (PBS)	Multi-vitamin tablets	N/A	[32]
Carbon paste	Phenol	Electrochemical oxidation NaOH	Tap and waste water	N/A	[33]
GC	Mn	$\text{NH}_3\text{-NH}_4\text{Cl}$	Lake water	0.001	[34]
Carbon film	Glucose	Electrochemical oxidation PBS	Wines	27	[35]
GC	Dopamine	Electrochemical oxidation PBS	Urine	0.008	[36]
SPE	Organophosphate	Covalent bond (electrochemical reduction of 4-nitrobenzene) and enzyme attachment		0.0001	[37]
GC	Erythromycin	Electrochemical oxidation $\text{NH}_3\text{-NH}_4$	Tablets	N/A	[41]
GC	Guanine, adenine	Covalent bond <i>r</i> -aminobenzene	Thiocholine	0.03	[42]
GC	Uric acid	Covalent bond $-\text{N}^+(\text{CH}_3)$	Urine	0.02	[43]
Carbon paste	Riboflavin	Covalent-bound aza macro-cycles	Pharmaceutical preparations and food	N/A	[45]

Table 3 Applications of polymer-modified electrodes

Polymer	Analyte	Modifier	Application	Detection limit (μM)	Ref
Polypyrrole	Nitrate	Nitrate, parylene	Potentiometric in situ nitrate biosensor	50	[53]
Poly(L-arginine)	Epinephrine	L-Arginine	Simultaneous determination	0.7	[54]
	Dopamine			0.3	
Melanin	Dopamine	N/A	N/A	0.005	[55]
L-Cysteine	Metal ions	N/A	N/A	N/A	[56]
Titania-sol-gel	Glucose	GOD/FcPF ₆	Serum glucose	N/A	[57] ^a
Eastman Kodak AQ55	X ⁻	Cyt C	Electrocatalysis studies	N/A	[58]
Nafion [®]	X ⁺	Fe ²⁺ , Fe ³⁺	N/A	<1	[59]
Tosflex	X ⁻	Fe(CN) ₆ ⁴⁻ , Fe(CN) ₆ ³⁻	N/A	<1	[59]
Nafion [®]	MPT	NiTSPc	Organophosphates	1–50 ng7 ml	[60] ^a
Methyl silicate		Antimony-doped tin oxide (ATO)	Transparent and conductive electrodes	N/A	[61]
<i>o</i> -Aminophenol	NADH	NADH detection	N/A	0.15	[65]
<i>o</i> -Phenylenediamine	Paracetamol	Copolymer aniline	N/A	1..5	[66]
Poly-toluidine blue	NADH	Electro-oxidative polymerisation	N/A	0.1	[70] ^a
L-Arginine	Epinephrine Dopamine	N/A	Human urine	2.93	[71]
Polyaniline	H ₃ O ⁺	N/A	pH measurements	N/A	[72]
Polyaniline		PB-oxidase	Oxidase biosensors	N/A	[73] ^a
Polyaniline	Choline	MWNTs	N/A	0.3	[74]
Polypyrrole	Ag ⁺	Sulphonated calixarenes	Ag ⁺	N/A	[75]
Pyrrole- <i>co</i> -pyrrole	Phenol	Amine and <i>p</i> -toluene sulphonated anions	Phenol	N/A	[76]
Polypyrrole	O ₂ ⁻	Anthraquinone(s)	Energy conversion systems	N/A	[77]
Polypyrrole	AA	FCN-Fe-Fe ₃ O ₄	AA in water samples	150	[78]
Pyrrole	H ₂ O ₂	HRPFe(CN) ₆ ⁴⁻ Fe(CN) ₆ ³⁻	Hydrogen peroxide biosensors	N/A	[79]
OPP	Catecholamines	Au nano-clusters	Human blood-serum	N/A	[83]
PQQ	Glucose	Glucose dehydrogenase	Independence of O ₂	N/A	[84]
Os-gel-HRP	Histamine	HAO	N/A	N/A	[85] ^a
Os-gel-HRP	Glucose, choline glutamate	GOD-ChOD-GlutaOD	N/A	0.01	[86]
Os-gel-HRP	Hypoxanthine	XOD	N/A	2	[87] ^a
Os-gel-HRP	Glutamate	GluDH-NOX	N/A	0.02	[88] ^a

^a Coupled with flow systems

state) and molecular filter properties. This has been shown to impart enhanced selectivity and sensitivity toward neurotransmitter analytes. They also confer the advantage of being able to detect those species that would not normally be electrochemically accessible. This relates to the exchange of the charged species within the cationic film (as counter ions) which alters the conductivity of the polymer backbone and thereby gives rise to the analytical signal [73].

Pre-formed redox gels consist of the other main polymer grouping and almost invariably incorporate osmium complexes [85–88]. The gel serves to both entrap the enzyme at the electrode surface and relay electrons from the carbon surface to the redox centre within the protein. The basis mechanism is highlighted in Fig. 3 using an example of peroxide reduction in which an osmium gel (Os-gel)–

horseradish peroxidase is employed. Hydrogen peroxide is reduced enzymatically by the immobilised enzyme which in turn is converted to the oxidised form. It is re-activated by the transfer of electrons from the electrode but which are transported via the osmium shuttles (Os(III)↔Os(II)). The

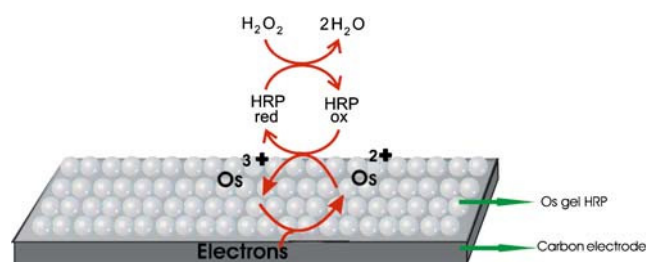


Fig. 3 Redox gel reaction scheme

Table 4 Carbon composite electrodes and their applications

Binder	Analyte	Modification	Applications	Detection limit (μM)	Ref
Nujol	Tryptophan	MWCNT–cobalt salophen	Human serum	0.1	[89]
Mineral oil	Cadmium	1-Furoylthioureas	Cadmium determination	60 $\mu\text{mg/l}$	[90]
Paraffin oil	Sugars and AA	Polyethylene glycol and Cu_2O	Biological samples	0.X	[93] ^a
Olive oil	Phenols	NA	Extra virgin olive oils	N/A	[94]
Paraffin	Phenol	Montmorillonite	Tap water	0.05	[95]
Paraffin oil	Anion	Ferrocene functionalised Calyx[4]pyrrole	Water	<0.1	[96]
Wax	Hydrazine	Manganese hexacyanoferrate	Photographic developer	6.65	[98] ^a

^a Coupled with flow systems

gels possess generic applicability allowing a host of enzyme systems to be used. A number of examples are highlighted in Table 3. The main drawback however is the fact that it requires the Os-gel to be cast onto the electrode surface [84–88].

Carbon composite electrodes

Composite electrodes typically consist of two or more components which together form the bulk of the base substrate. The composition of this group can be incredibly diverse depending on definition but, in general, the field can be sub-divided into carbon paste [89–109] and screen-printed [110–118] systems. Despite the different formats and processing requirements—both share a common methodology in which carbon particles are bound together by a binder. The great strength of this approach is the fact that the composition—and hence properties—of the electrode can be easily manipulated by altering one or more of the components in the paste or ink or by the simple addition of a catalyst or other modifier during the initial mixing phase. Both systems have been used in LC applications and each approach can confer particular operational advantages.

Carbon paste electrodes typically rely upon carbon particles being held together by an impregnated organic liquid phase that is compacted into a holder and can then be

used in much the same way as a conventional solid electrode and likewise can be incorporated within flow systems in much the same way as a commercial glassy carbon or HOPG electrode [89–98]. Their application within such is however less common—as indicated in Table 4. This is rather surprising as—irrespective of the particular form of carbon used—the paste electrodes can be rapidly constructed with little expense beyond the initial cost of the components. As such, they can facilitate a prototyping function that allows the ad hoc production of modified electrodes. Operationally, the paste electrodes can have the advantage of providing a low background compared to solid graphite and noble metal electrodes. They are also renewable in that, upon fouling, the surface can be regenerated by polishing [89, 93]. There is clearly the potential for a huge number of variations in the preparation—the form of carbon used, the nature of the binder and presence of modifiers or catalysts. The binder can take the form of a simple mineral oil (nujol or paraffin) [89–98], wax or epoxy [98], polymers [67] or ionic liquid [99].

CNTs have risen to considerable prominence in recent years and have been widely investigated (Table 5) as an electrode modifier or as a direct replacement for the more traditional graphitic powder [100–109]. The duality of construction emphasises the versatility of the material.

Table 5 Carbon nanotubes electrodes and their applications

Analyte	Modification	Applications	Detection limit (μM)	Ref
Nitrite, AA, uric acid	Polypyrrole	N/A	N/A	[72]
Homocysteine	Nafion [®] , CNT	N/A	0.06	[103] ^a
Galactose	Nafion [®] /CNT	Carbohydrate detection	10	[104] ^a
Chlorate, iodate, bromate	MWCNT–iron (III)–porphyrin	Chlorate, iodate, bromate	0.5, 2, 0.001	[105]
Estradiol	Nafion [®] /Ni (Cyclam)	Serum	0.06	[106]
Glucose	PtPbNP	Blood	N/A	[107]
H_2O_2	Nafion [®] –Au–GOD	Beverages	17	[108]
NH_3	SiO_2 –Si–Li	N/A	N/A	[109]

^a Coupled with flow systems

Table 6 Screen-printed electrodes and their applications

Analyte	Modification	Applications	Detection limit (μM)	Ref
H_2O_2	HRP	N/A	N/A	[111] ^a
Organophosphates	Tyrosinase	River water	N/A	[112]
Nitric oxide	Direct oxidation	Biological samples	N/A	[110] ^a
H_2O_2	GOD	Glucose	N/A	[107]
Glucose	Osmium complex	N/A	N/A	[114] ^a
AA	$\text{Fe}(\text{CN})_6^{4-}$, $\text{Fe}(\text{CN})_6^{3-}$	N/A	N/A	[115] ^a
IgG	Antigen–antibody	Immuno-enzyme	3 ng/ml	[116] ^a
Sulphide	Cinder/hexacyano cobaltate	N/A	N/A	[117] ^a
H_2O_2	Copper	N/A	0.99	[118] ^a

^a Coupled with flow systems

The nanotubes can generally be found in two forms: single wall (SWCNT) or multi-walled (MWCNT) and, by analogy to graphite, the walls of the nanotube can be defined as basal plane while the ends are considered to be edge plane. It has been demonstrated that the edge plane effects, often in combination with increased oxygen functionalities, are largely responsible for the enhanced electrode kinetics [28, 30]. Reviews covering the breadth of nanotube application and the fundamentals that underpin their action are widely available. In common with the paste systems, they have the notable benefit of decreasing the capacitive background which increases the sensitivity of the electrode [103]. As in the case of the carbon paste systems, the exploitation of nanotubes within flow systems is more limited despite the considerable advantages they can potentially confer. This is likely to be a consequence of the field being relatively young and it could be expected that more electrode designs incorporating nanotubes will emerge in the future.

Screen-printed electrodes are generally inexpensive to fabricate when considering large-scale processing in which the costs of individual electrodes can be expected to fall dramatically. Access to such facilities and the large volumes of the reagents required can however be prohibitive—especially when considering the incorporation of expensive biological reagents (cf. carbon paste systems). Mass manufacturing does however open an avenue to the production of disposable electrodes that can overcome the fouling and surface deactivation effects that repeated analysis can bring. The batch production of near identical

sensors can therefore minimise the reproducibility problems associated with the regeneration of conventional electrochemical detectors [110–118]. Some of the applications are detailed in Table 6

The screen-printed electrodes can be integrated into conventional flow systems—operating under planar or the wall jet schemes outlined in Fig. 1. The latter method is highlighted in the analysis of hydrogen peroxide on a copper-modified screen-printed carbon electrode [118]. The radial flow of the analyte across the wall of the electrode minimises the diffusion layer which amplifies the signal and can give a significant advantage over the planar arrangement.

Diamond electrodes

The increasing availability of diamond electrodes hold much promise for electroanalytical applications as they can offer substantial advantages over conventional glassy carbon or HOPG systems [119–123]. They generally possess a much wider electrochemical window in aqueous and non-aqueous media. It has also been shown that polar molecules adsorb less at the surface and, as such, the electrode operating lifetime can be significantly greater than that achievable with conventional carbon surfaces. It has been shown that the repeated oxidation of catechol can be achieved without the cumulative loss in sensitivity that occurs in biological matrices [119]. Boron-doped diamond also demonstrates a marked insensitivity to dissolved

Table 7 Applications of diamond electrodes in flow systems

Analyte	Applications	Detection limit (μM)	Ref
Azide ions	Azide detection	0.004	[119]
Dopamine (in presence of AA)	Dopamine detection	0.0025	[119]
Histamine and sulphadiazine	Histamine and sulphadiazine detection	N/A	[120]
Chlorophenols and theophylline	Theophylline and chlorophenols detection	0.5	[122]

oxygen and thereby provides a significant improvement over glassy carbon or HOPG when considering cathodic operations [120]. A range of applications in which the performance of boron-doped diamond electrodes has been investigated are compared and contrasted in Table 7.

Future directions

It is clear that carbon possesses considerable flexibility in its ability to be modified and effectively tailored to specific applications and stands in contrast to the more limited options available when considering metallic electrodes. These features have been long recognised for general electroanalytical investigations but are now increasingly finding application as detectors in flow systems. While bare carbon substrates were once the default option in electrochemical detection, the modified systems are capable of significantly enhancing performance. This has been demonstrated by the use of diamond and nanotube composite electrodes which can extend the range of accessible analytes and the sensitivity of the resulting response. The requirement for miniaturised systems has seen the introduction of carbon within micro-analytical systems and lab-on-a-chip prototypes [87, 98, 97].

Dual electrode systems have long been used in flow systems and have been reported to offer considerable advantages over conventional single-electrode configuration [124–127] detailed in Fig. 1. The couple essentially acts as generator–collector systems in which it can offer two advantages—the product of the first electrode process may provide a more easily quantifiable signal than the parent compound (i.e. it may be detected at a potential where there are few interference—especially in complex matrices where co-elution of species may be a factor) [128, 129]. Depending on the arrangement within the cell—the electrode can also be set up to induce repeated redox cycling between oxidised and reduced states and thereby offer tremendous gains in sensitivity [126–130]. These early designs were largely developed using metal electrode systems [124–127] but carbon has steadily made inroads to this area [128–134] and the use of micro-arrays is a future addition in which the main gain is the possibility of multi-parametric detection within a single sample.

Issues over fouling and the consequent compromises that occur in terms of reproducibility could be countered through the adoption of screen-printed detectors. It could be envisaged that these could be integrated within conventional LC architectures but, more importantly, their inherent disposability could herald an opportunity for the fabrication of stand-alone miniaturised chip–sensor-based systems. These could eventually induce the transfer of the more traditional laboratory-based LC to decentral-

ised testing environments. A diverse range of applications is apparent from the various tables and it can be seen that carbon, as the core sensing detection element, can have a significant impact on a highly varied range of chemicals that cross the agri-food, biomedical, environmental and industrial sectors.

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